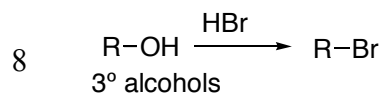
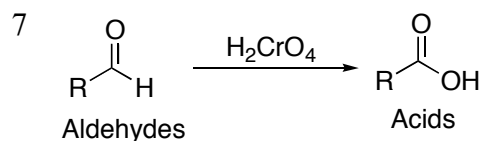
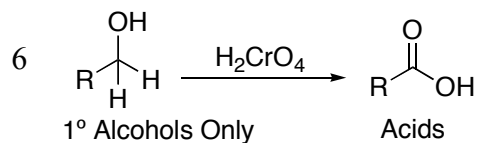
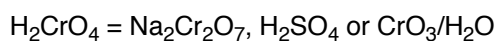
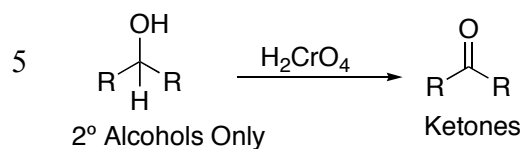
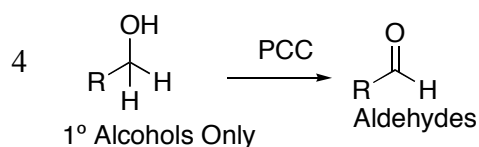
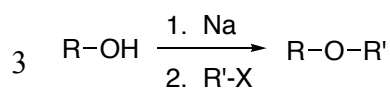
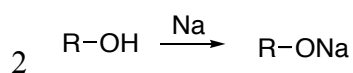
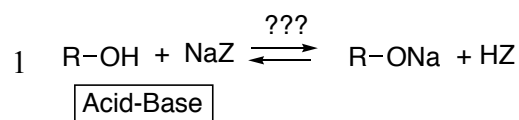


Summary of Alcohol Reactions, Ch. 11.

Mech: Be able to draw!

1. Deprotonation by a base.
2. Controlled by relative stability of RO^- versus Z^- .
3. Consider relative electronegativity and whether either anion is resonance stabilized.

- Potassium (K) analogous.
- Key way to convert alcohol to alkoxide, reactive as $\text{S}_{\text{N}}2$ nucleophile and E2 base.

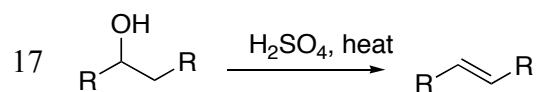
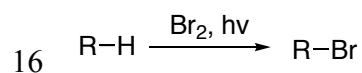
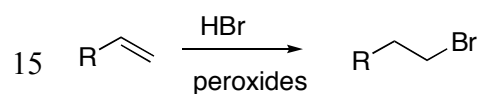
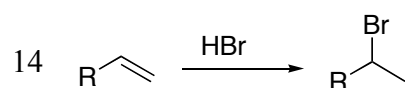
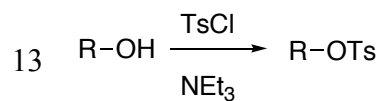
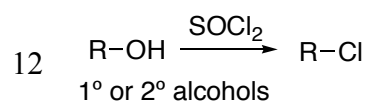
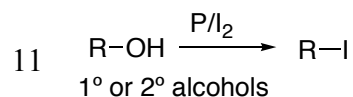
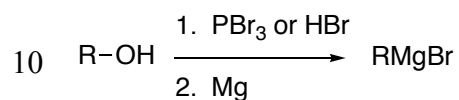
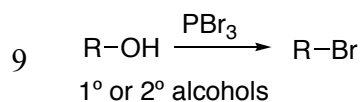
- Alkoxide formation- $\text{S}_{\text{N}}2$ route to ether
- The electrophile $\text{R}'\text{-X}$ must be $\text{S}_{\text{N}}2$ reactive, preferably 1° with a good leaving group

- Key access to aldehydes, which are useful for more Grignard chemistry.
- Note difference between PCC and H_2CrO_4
- PCC does not react with 2° alcohols very rapidly

- Key access to ketones.
- PCC does not react very fast with 2° alcohols

- Note difference between PCC and H_2CrO_4 when reacting with 1° alcohols.

- HI, HCl analogous
- Converts alcohol into a bromide that can be used in Grignards, E2 reactions
- Cation mechanism
- Usually not method of choice for $1^\circ, 2^\circ$ alcohols



- Converts alcohol into a bromide that can be used in Grignards, E2, S_N2 reactions
- Inversion of stereochem
- Not good for 3° alcohols

- Quick 2-step conversion of alcohol into a nucleophilic Grignard

- Via PI₃

- Retention of stereo!

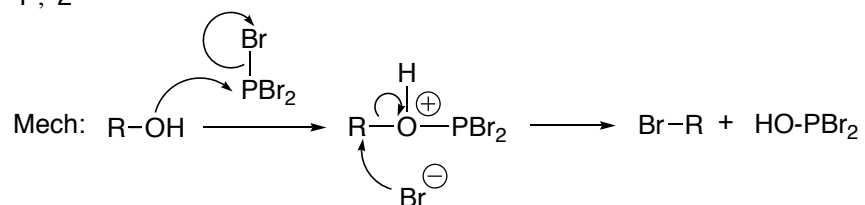
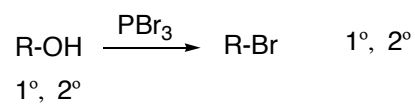
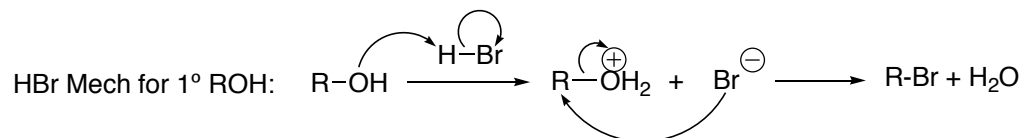
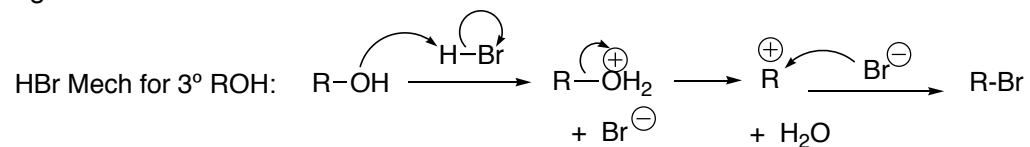
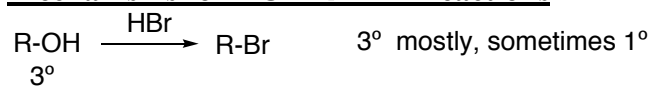
- Tosylates are super leaving groups, better even than iodides.
- Tosylates are well suited to S_N2 and E2 reactions.

- Markovnikov addition

- anti-Markovnikov addition

- Radical mechanism, 3° > 2° > 1°

- Zaytsev elimination

Mechanisms for ROH \rightarrow RBr Reactions

Ch. 11 Reactions of Alcohols

A. Conversion to Alkoxides (Sections 11.14, 10.6)

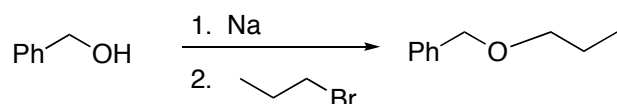
“alkoxide” = RO^- anion

- By acid-base deprotonation (Section 10.6)
 - A rather reactive anion base is required that is *less* stable than an alkoxide anion
 - Carbanions (RMgBr) or nitrogen anions can do this
 - NaOH can't
- By redox reaction with sodium or potassium (or some other metals)

1	$\text{R-OH} + \text{NaZ} \xrightleftharpoons{???} \text{R-ONa} + \text{HZ}$ <div style="border: 1px solid black; display: inline-block; padding: 2px;">Acid-Base</div>	<ol style="list-style-type: none"> Deprotonation by a base. Controlled by relative stability of RO^- versus Z^-. Consider relative electronegativity and whether either anion is resonance stabilized.
2	$\text{R-OH} \xrightarrow{\text{Na}} \text{R-ONa}$	<ul style="list-style-type: none"> Potassium (K) analogous. Key way to convert alcohol to alkoxide, reactive as $\text{S}_{\text{N}}2$ nucleophile and E2 base.

B. Conversion to Ethers via Alkoxide (11-14)

3	$\text{R-OH} \xrightarrow[2. \text{R}'\text{-X}]{1. \text{Na}} \text{R-O-R}'$	<ul style="list-style-type: none"> Alkoxide formation-$\text{S}_{\text{N}}2$ route to ether The electrophile $\text{R}'\text{-X}$ must be $\text{S}_{\text{N}}2$ reactive, preferably 1° with a good leaving group
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C. Oxidation of Alcohols to Carbonyl Compounds (11.1-4)

Summary: 2 Oxidants

1. **PCC = mild** **1° alcohols \rightarrow aldehydes**

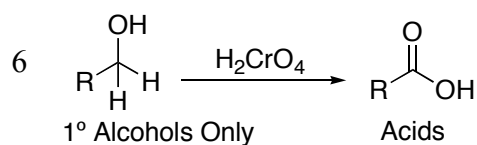
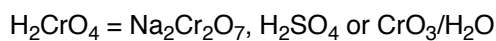
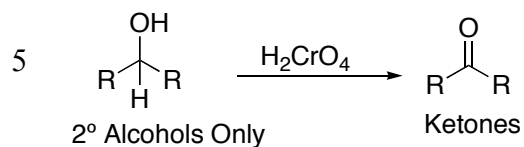
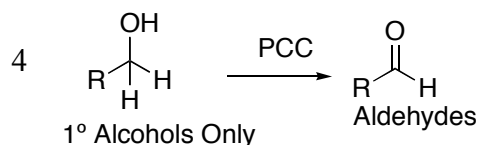
- “**P**yr**Y**ridinium **ch**loro**ch**romate”: soluble in water-free dichloromethane
- Mild, selective for 1° over 2° alcohols, and when 1° alcohols are used stops at aldehyde

2. **H_2CrO_4 = strong**a. **2° alcohols \rightarrow ketones**b. **1° alcohols \rightarrow carboxylic acids**c. **3° alcohols \rightarrow no reaction**d. **aldehydes \rightarrow carboxylic acids**

- $\text{H}_2\text{CrO}_4 = \text{CrO}_3 + \text{H}_2\text{O}$ or $\text{Na}_2\text{Cr}_2\text{O}_7 + \text{H}_2\text{SO}_4$ (make in the reaction flask)

- Always made and used in the presence of some water

- Very strong, when 1° alcohols are used goes $1^\circ \text{RCH}_2\text{OH} \rightarrow \text{RCHO} \rightarrow \text{RCO}_2\text{H}$ without stopping at aldehyde

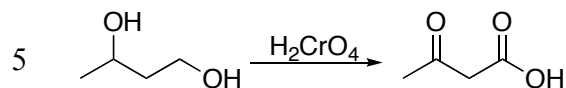
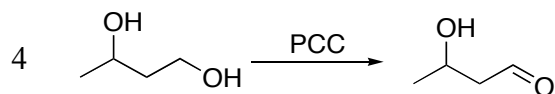
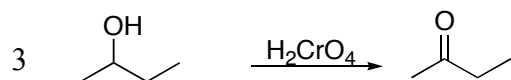
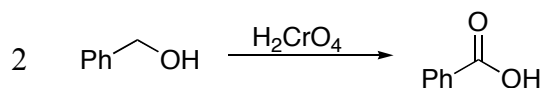
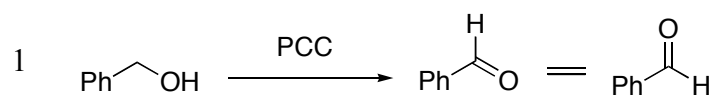


- Key access to aldehydes, which are useful for more Grignard chemistry.
- Note difference between PCC and H_2CrO_4
- PCC does not react with 2° alcohols very rapidly

- Key access to ketones.
- PCC does not react very fast with 2° alcohols

- Note difference between PCC and H_2CrO_4 when reacting with 1° alcohols.

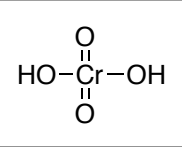
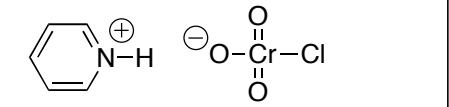
Draw the products for the following oxidation reactions.

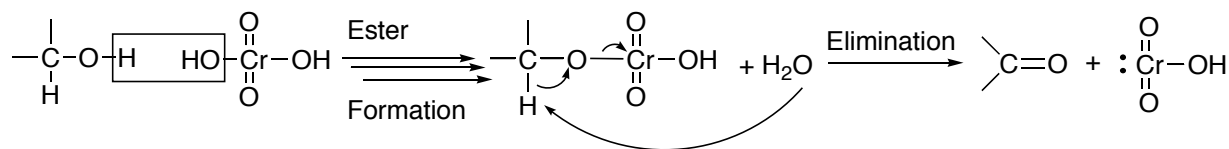


Jones Test H_2CrO_4 for Alcohols (11-2C) (test responsible)

- H_2CrO_4 (Jones Reagent) is clear orange
- Treatment of an unknown with Jones reagent:
 - Solution stays clear orange \rightarrow no 1° or 2° alcohol present (negative reaction)
 - Solution gives a green/brown precipitate \rightarrow 1° or 2° alcohol present (positive reaction)
 - 3° , vinyl, and aryl alcohols do not react. Nor do ketones, ethers, or esters.

Structure and Mechanism (not test responsible)

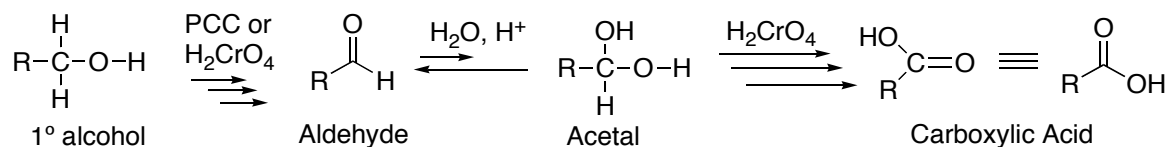
$\text{H}_2\text{CrO}_4 = \text{chromic acid} = \text{Na}_2\text{Cr}_2\text{O}_7 = \text{CrO}_3/\text{H}_2\text{O} = \text{Cr}^{+6}$ oxidation state <ul style="list-style-type: none"> • Water soluble 	
Pyridinium carbons renders PCC soluble in organic solvents, thus it is functional in organic solvent and in the absence of water	 <p>PCC = Pyridinium ChloroChromate</p>

General Mechanism

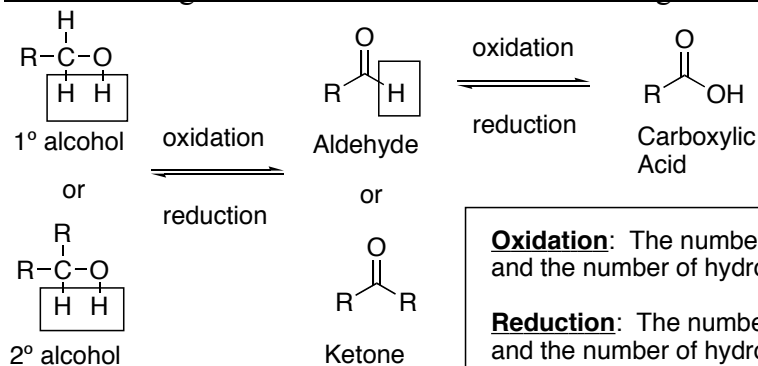
- PCC operates analogously

 1° Alcohols, Aldehydes, and the Presence or Absence of Water:

Q: Why does Anhydrous PCC stop at Aldehyde but Aqueous H_2CrO_4 Continues to Carboxylic Acid?



1. Both PCC and H_2CrO_4 convert 1° alcohols to aldehydes
2. In the presence of acidic water, aldehydes undergo an equilibrium addition of water to provide a small equilibrium population of acetal
3. The acetal form gets oxidized (very rapidly) to carboxylic acid
 - The aldehyde form cannot itself get oxidized to carboxylic acid
 - Since PCC is used in absence of water, the aldehyde is not able to equilibrate with acetal and simply stays aldehyde.
 - Since it can't convert to acetal, therefore no oxidation to carboxylic acid can occur
4. Chromic acid, by contrast, is in water
 - Therefore the aldehyde is able to equilibrate with acetal
 - The acetal is able to be oxidized.
 - Thus, the aldehyde via the acetal is able to be indirectly oxidized to carboxylic acid, and in fact does so very rapidly.

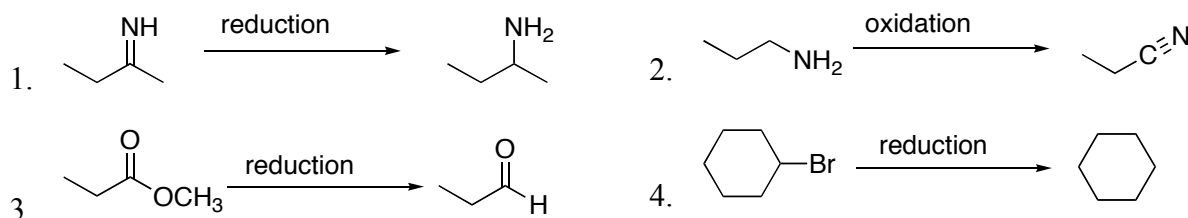
General Recognition of Oxidation/Reduction in Organic Chemistry

Oxidation: The number of oxygen bonds to a carbon increases, and the number of hydrogens bonded to a carbon decreases

Reduction: The number of oxygen bonds to a carbon is reduced, and the number of hydrogens bonded to a carbon increases.

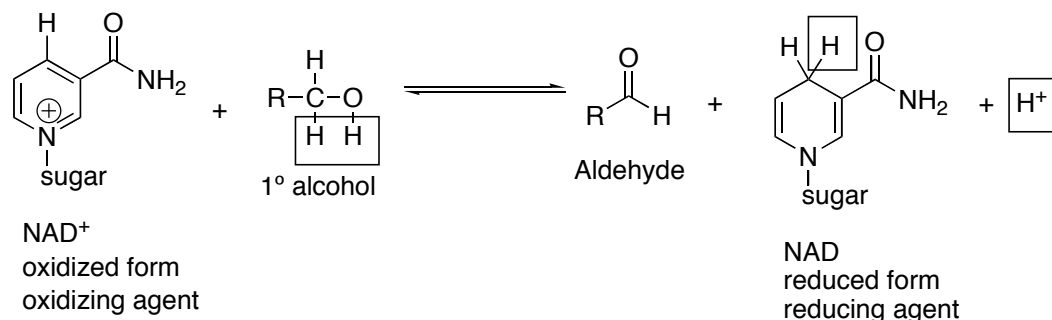
More General: # of bonds to heteroatoms versus to hydrogens

Classify the following transformations as “oxidations” or “reductions”

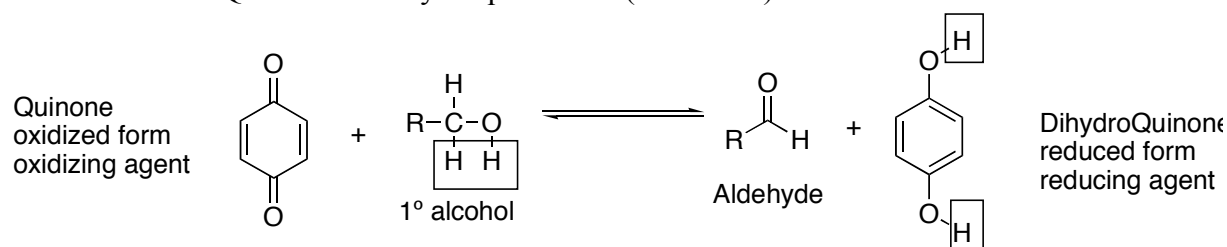
11.3, 11.4 Other methods for Oxidizing Alcohols. (No test)

There are lots of other recipes used for oxidizing alcohols (and for other oxidation reactions)

1. KMnO_4
2. CuO
3. “Jones”: H_2CrO_4 with acetone added to temper reactivity
4. Collins: H_2CrO_4 with pyridine added to temper reactivity
5. “Swern”: $(\text{COCl})_2$ and $(\text{CH}_3)_2\text{S}=\text{O}$ then NEt_3
6. HNO_3
7. Biological Oxidant 1: “ NAD^+ ” “nicotinamide adenine dinucleotide”



8. Biological Oxidant 2: “Quinones and hydroquinones” (Ch. 17-15)



In General: Recognizing Oxidizing versus Reducing Agents**Oxidizing Agents:** Often have:

- Highly Oxidized Metals or Nonmetals
- Extra Oxygen

OsO₄ (+8)
 KMnO₄ (+7)
 CrO₄ (+6)
 H₂CrO₄ (+6)
 HNO₄ (+5)
 H₂O₂ → H₂O
 RCO₃H → RCO₂H
 O₃ → O₂

Reducing Agents: Often involve:

- Hydrides in Formulas
- Highly Reduced Metals
- Metals + H₂
- Metals + acid

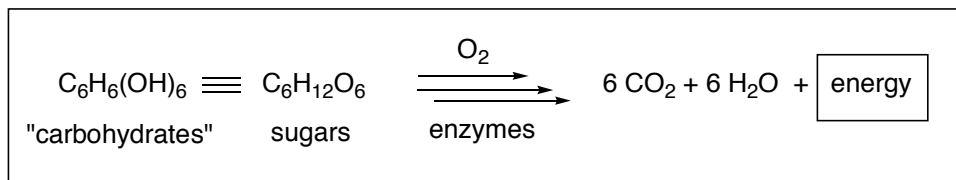
LiAlH₄
 NaBH₄
 Li, Na, K, Mg, Zn, Al, etc.
 Pd/H₂, Pt/H₂, Ni/H₂ etc.
 Zn/HCl, Fe/HCl, Zn/Hg/HCl, etc..

- The ability to qualitatively recognize when a transformation involves an oxidation or reduction can be very helpful.
- The ability to recognize a reactant as an oxidizing agent or a reducing agent can be very helpful
- Often on standardized tests!

Some Biological Alcohol Oxidations (Not for Test)

1. Oxidation of “carbohydrates” or “sugars” is the primary source of bioenergy

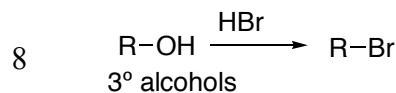
- multiple enzymes are involved for the many steps
- A “carbohydrate” basically has a formula with one OH per carbon



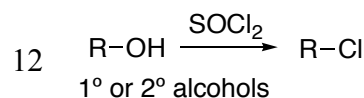
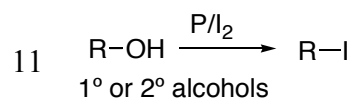
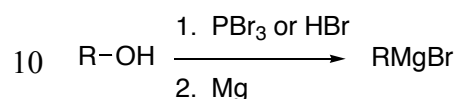
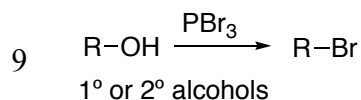
2. Most alcohols are biooxidized to give toxic carbonyl derivatives (“intoxication”)

- the presence of substantial aldehydes and especially ketones in the blood is symptomatic of various problems
 - intoxication
 - alcoholism
 - uncontrolled diabetes
 - etc (other metabolic disorders)

11.7-9 Conversion of Alcohols to Alkyl Halides



Mech: Be able to draw!



- HI, HCl analogous
- Converts alcohol into a bromide that can be used in Grignards, E2 reactions
- Cation mechanism
- Usually not method of choice for 1°, 2° alcohols

- Converts alcohol into a bromide that can be used in Grignards, E2, S_N2 reactions
- **Inversion of stereochem**
- Not good for 3° alcohols

- Quick 2-step conversion of alcohol into a nucleophilic Grignard

- Via PI₃

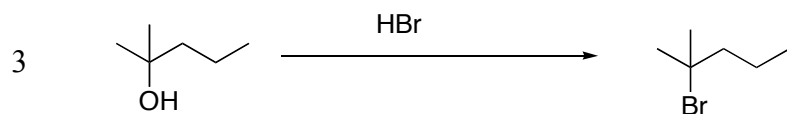
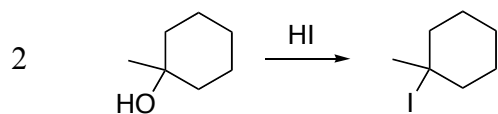
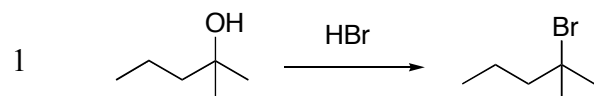
- Retention of stereo!
- Section 11-9

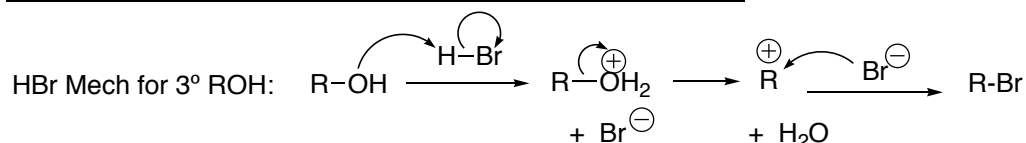
Summary:

Class	<u>R-Br</u>	<u>R-I</u>	<u>R-Cl</u>
1° ROH	PBr ₃	P/I ₂	SOCl ₂
2° ROH	PBr ₃	P/I ₂	SOCl ₂
3° ROH	HBr	HI	HCl
Vinyl or Aryl	Nothing works	Nothing works	Nothing works

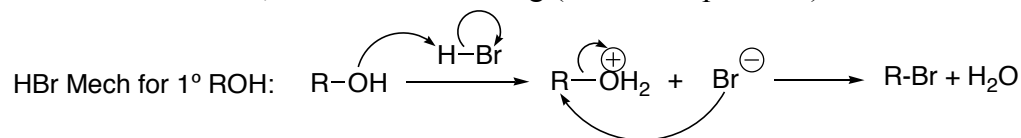
Straight Reaction with H-X (Section 11.7)

- Ideal only for 3° ROH,
- sometimes works with 1° alcohols, with a complex mechanism
- Only occasionally for 2° alcohols
- **Method of choice for 3°, but not for 1° or 2°**



Mechanism for H-X reactions with 3° Alcohols: CationicNotes:

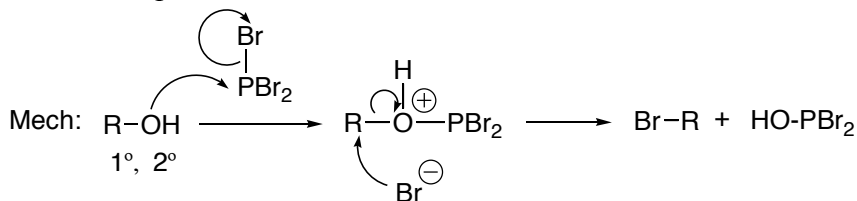
- Memorize the 3° alcohol mechanism (test responsible)
 - Protonate
 - Leave to give Cation. This is the slow step for 3° alcohols
 - Capture
- Analogous with HI or HCl
 - HCl slower, normally enhanced with ZnCl₂, which enhances rate of cation formation (Lucas test, see later)
 - Outside of 3° systems, side reactions are common and yields aren't often very good
- Outside of 3° alcohols, side reactions are common and yields aren't often very good
 - Elimination reactions and cation rearrangements...
- S_N1 type: carbocation-forming step is the rate-determining step, so R⁺ stability key
 - 3° alcohols fastest
 - 2° alcohols are way slower
 - 1° alcohols can't react at all via this mechanism, because 1° R⁺ are too unstable.
 - Ditto for vinyl or aryl alcohols
- HBr can also react with 1° ROH to give 1° RBr, although it is not often the method of choice
 - The mechanism is different, but rather interesting (not test responsible)



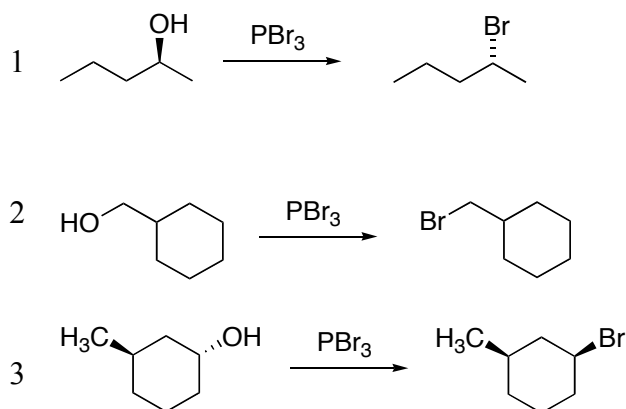
- carbocation formation never occurs
- bromide ion simply does S_N2 on the protonated alcohol, with water as an excellent leaving group
- yields tend to be pretty inconsistent

Reaction of 1° and 2° Alcohols with PBr₃ (Section 11-8)

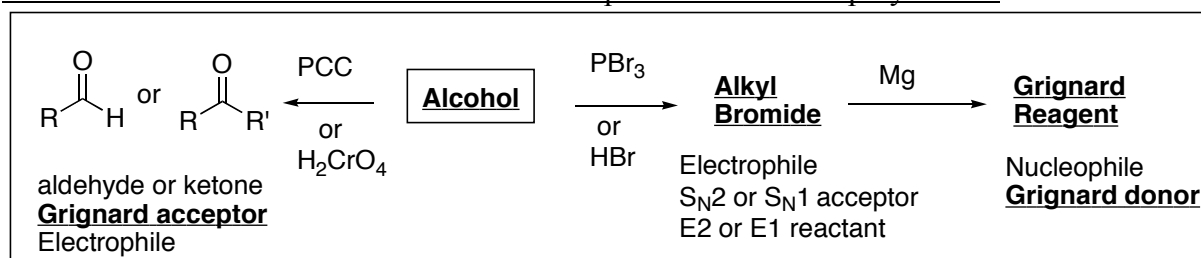
- Default recipe for 1° and 2° alcohols



- PBr₃ is an exceptional electrophile, and reacts even with neutral alcohols
- The first step activates the oxygen as a leaving group.
- The second step involves an S_N2 substitution
 - stereochemical inversion occurs chirality present (common for 2° alcohols)**
- Because the second step is an S_N2 substitution, the reaction fails for 3° ROH
- PCl₃ does not react as well, and is not useful for making chlorides
- PI₃ is not stable and can't be stored in a bottle. However, the combination of 1P + 1.5 I₂ → PI₃ in the reaction container (*in situ*)
 - Thus P/I₂ essentially provides the PI₃ that does the job



Conversions of Alcohols into Other Reactive Species in Multi-Step Syntheses



- oxidation can convert an alcohol into a carbonyl = **Grignard acceptor (electrophile)**
- PBr₃/Mg or HBr/Mg can convert an alcohol into RMgBr = **Grignard donor (nucleophile)**
- PBr₃ or HBr can convert an alcohol into RBr, capable of normal substitution and elimination reactions.

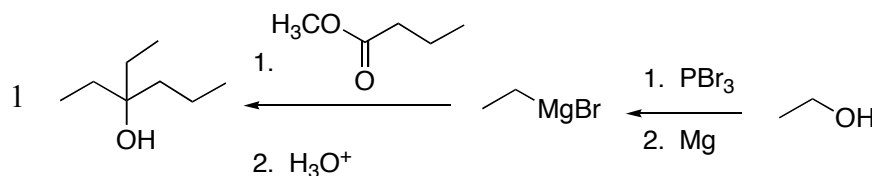
Retrosynthesis Problems (In which you decide what to start from): Design syntheses for the following.

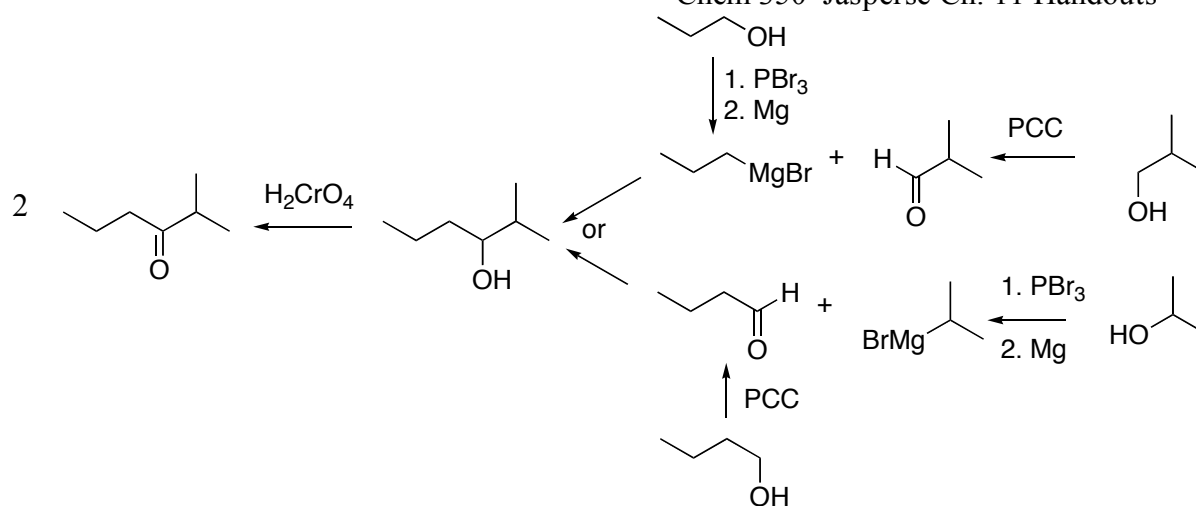
Allowed starting materials include:

Bromobenzene	cyclopentanol	any acyclic alcohol or alkene with ≤4 carbons
any esters	ethylene oxide	formaldehyde (CH ₂ O)
any "inorganic" agents (things that won't contribute carbons to your skeleton)		

Tips:

- Focus on the functionalized carbon(s)
- Try to figure out which groups of the skeleton began together, and where new C-C bonds will have been formed
- When "breaking" it up into sub-chunks, try to make the pieces as large as possible (4 carbon max, in this case, for acyclic pieces)
- Remember which direction is the "true" laboratory direction.
- Be careful that you aren't adding or subtracting carbons by mistake

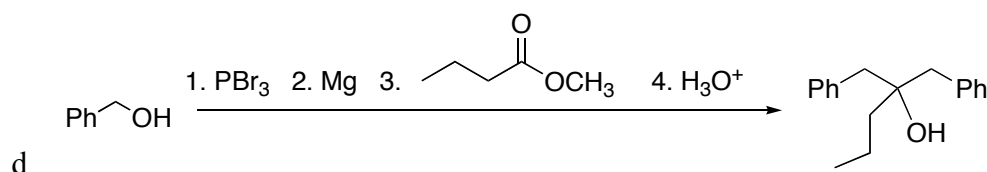
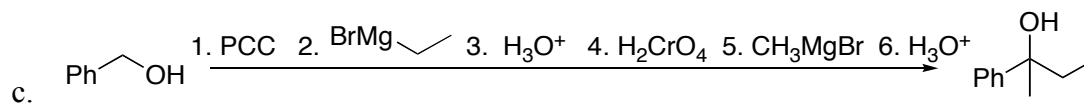
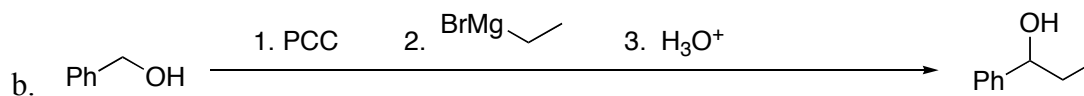
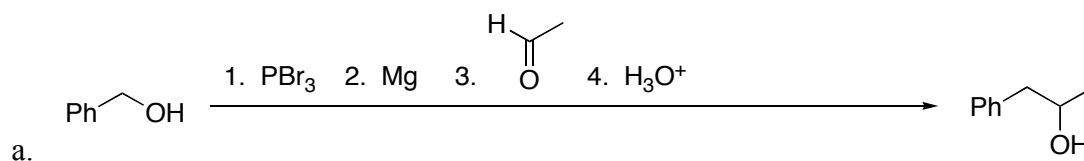


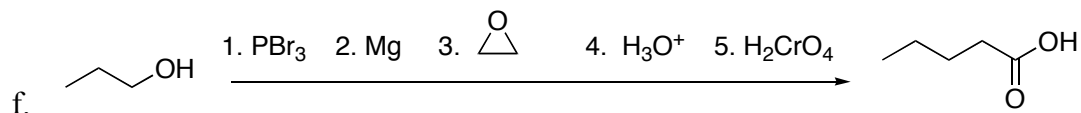
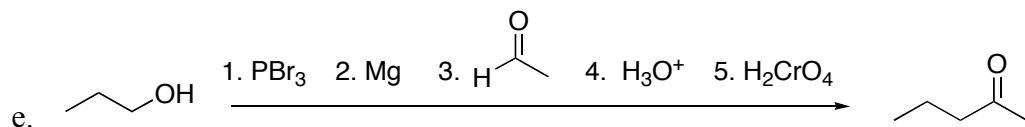


Normal Synthesis Design: In which you are given at least one of the starting Chemicals.

Provide Reagents. You may use whatever reagents, including ketones or aldehydes or Grignards or esters, that you need. Tips:

- Identify where the reactant carbons are in the product
- Is the original carbon still oxygenated? \rightarrow it will probably function as a Grignard acceptor
- Is the original carbon not still oxygenated? \rightarrow it should probably function as Grignard donor
- Working backwards helps.





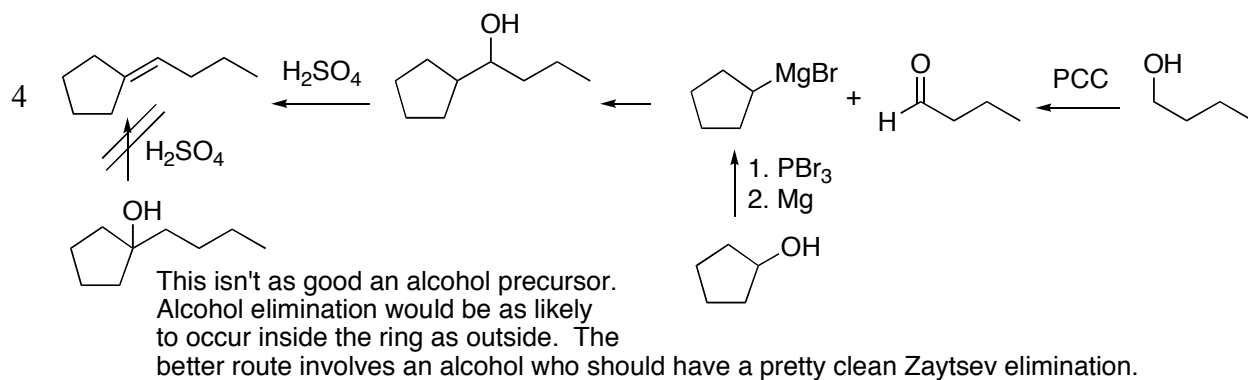
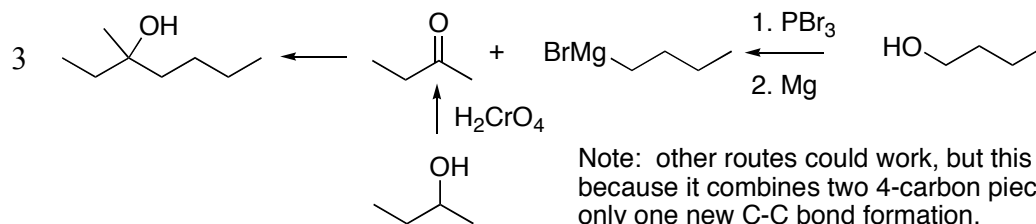
Retrosynthesis Problems: Design syntheses for the following.

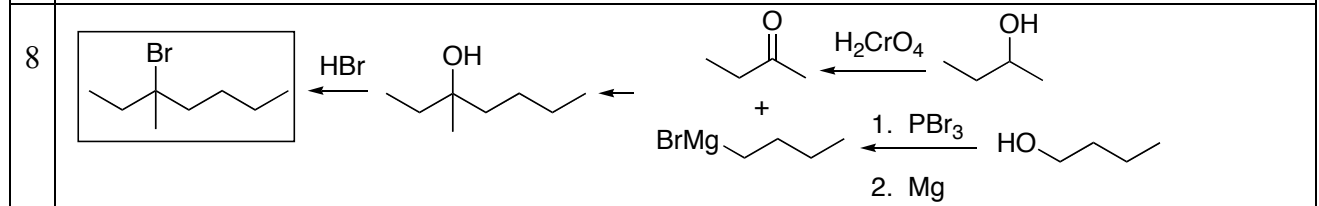
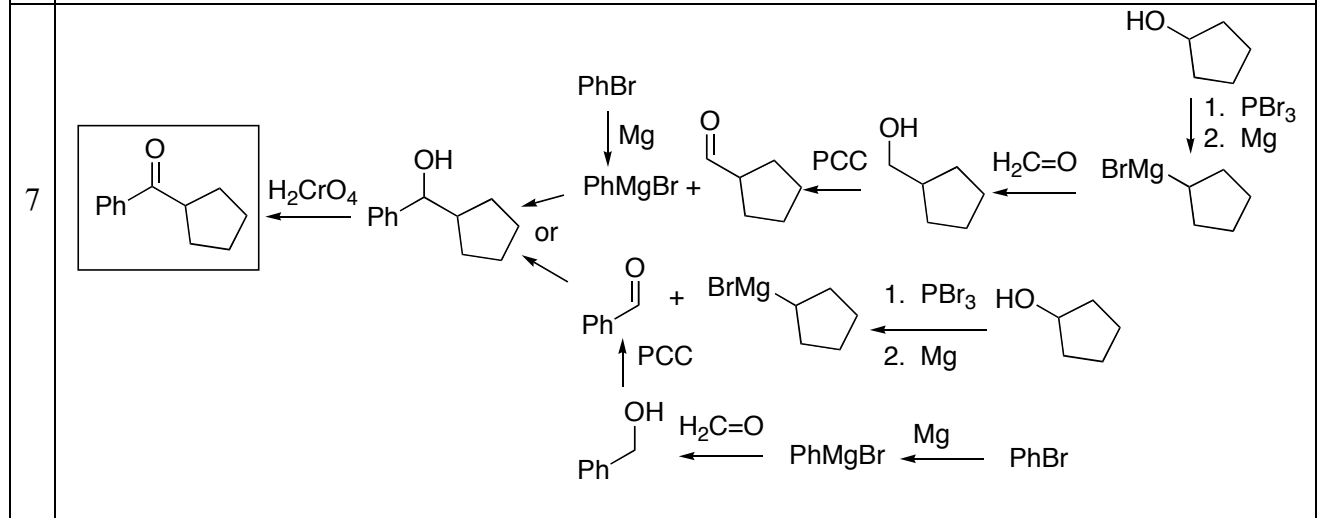
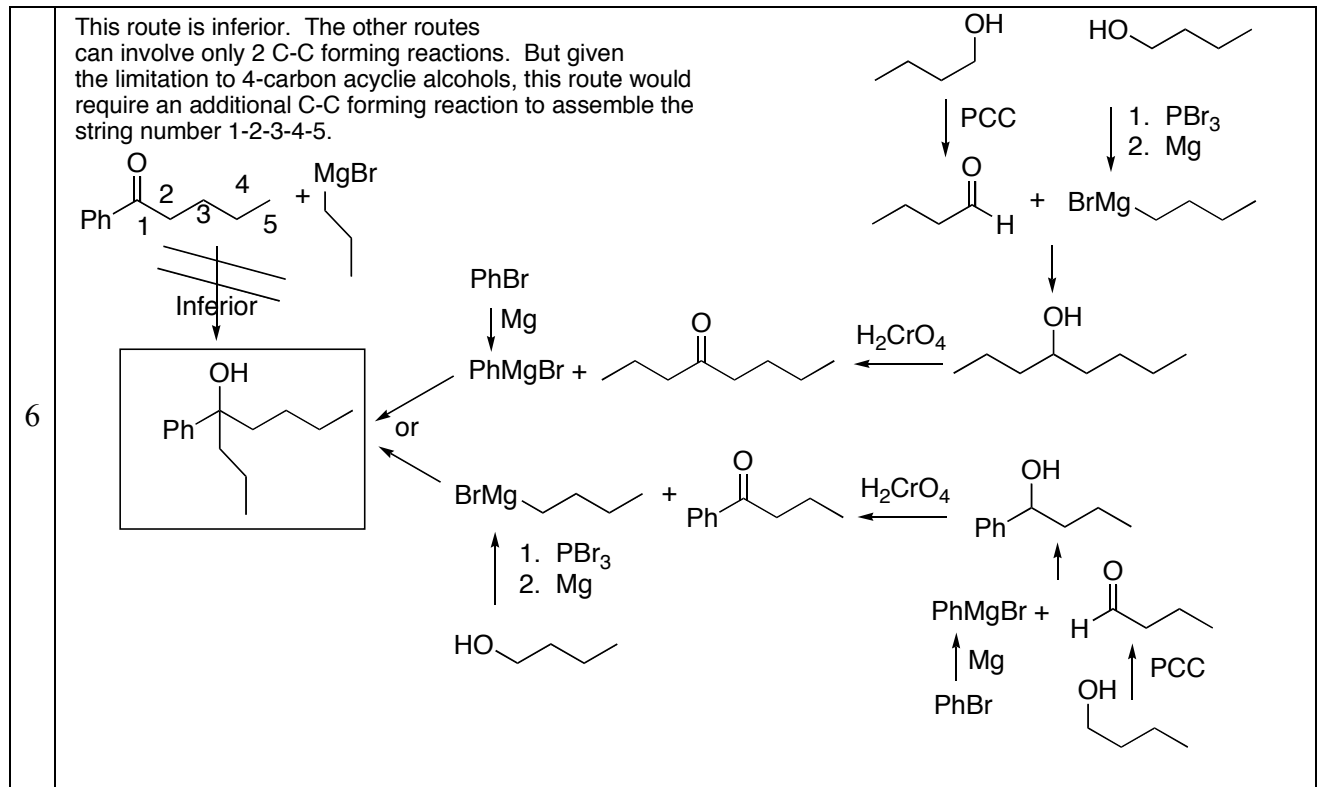
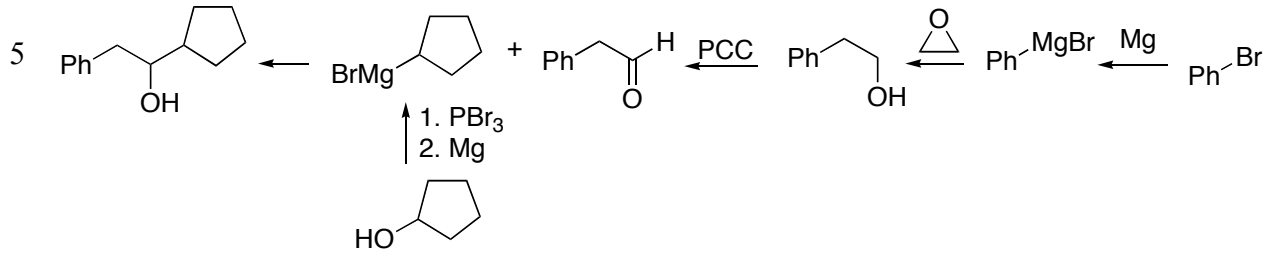
Allowed starting materials include:

Bromobenzene	cyclopentanol	any acyclic alcohol or alkene with ≤ 4 carbons
any esters	ethylene oxide	formaldehyde (CH ₂ O)
any "inorganic" agents (things that won't contribute carbons to your skeleton)		

Tips:

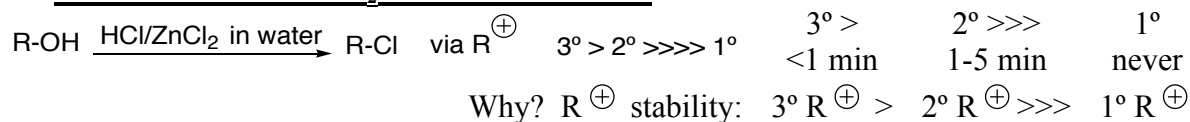
1. Focus on the functionalized carbon(s)
2. Try to figure out which groups of the skeleton began together, and where new C-C bonds will have been formed
3. When "breaking" it up into sub-chunks, try to make the pieces as large as possible (4 carbon max, in this case, for acyclic pieces)
4. Remember which direction is the "true" laboratory direction.
5. Be careful that you aren't adding or subtracting carbons by mistake



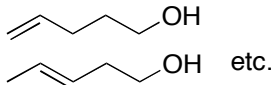
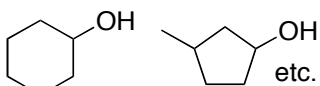
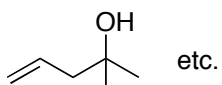
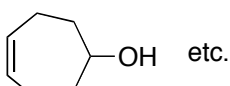
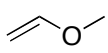
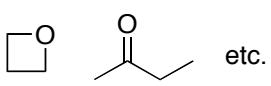
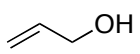
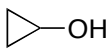
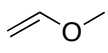


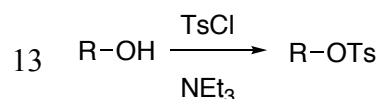
Unknowns and Chemical Tests (Sections 11-2C, 11-7)

- H₂/Pt test for alkenes
- Br₂ test for alkenes
- Jones reagent (H₂CrO₄) Test for 1° or 2° alcohols**
 - 3° alcohols do not react
 - 2° alcohols keep the same number of oxygens but lose two hydrogens in the formula
 - 1° alcohols lose two H's but also add one oxygen

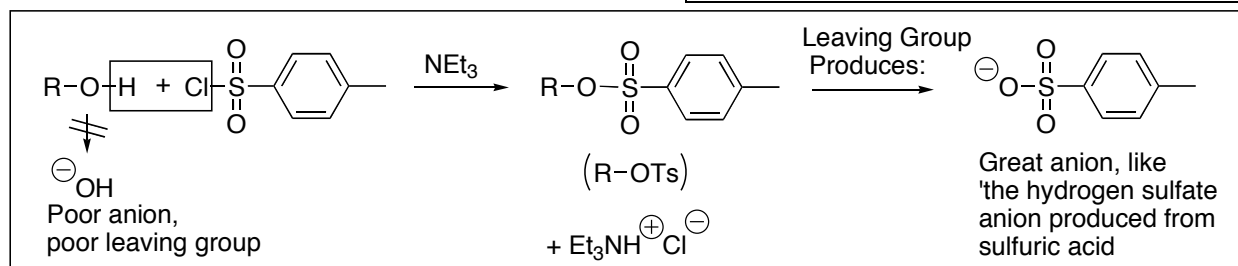
4. Lucas Test: HCl/ZnCl₂ for 3° or 2° alcohols

- 3° alcohols are fastest
- 1° alcohols don't react at all
- R[⊕] stability is the key
- Test is based on **solubility**: The R-Cl product is nonpolar and water insoluble, so it separates out from water. Alcohols are quite soluble especially in highly acidic water.
- Test fails is useless for alcohols with so many carbons that it doesn't even dissolve in the original HCl/ZnCl₂/water solution

	Jones (H₂CrO₄)	Lucas (HCl/ZnCl₂)	H₂/Pt	Required Facts	Possible Answers
1 C ₅ H ₁₀ O	Yes	No	Yes	1° ROH One Alkene	 etc.
2 C ₆ H ₁₂ O	Yes	Yes, 1-5 min	No	2° ROH One ring	 etc.
3 C ₆ H ₁₂ O	No	Yes	Yes	3° ROH One Alkene	 etc.
4 C ₇ H ₁₂ O	Yes	Yes	Yes, Produces C ₇ H ₁₄ O	2° ROH One Alkene One ring or carbonyl	 etc.
5 C ₃ H ₆ O	No	No	Yes	No ROH One alkene	
6 C ₃ H ₆ O	No	No	No	No ROH One ring or Carbonyl	 etc.
7 C ₃ H ₆ O	Yes, produces C ₃ H ₄ O ₂	No	Yes	1° ROH One alkene	
8 C ₃ H ₆ O	Yes, produces C ₃ H ₄ O	Yes	No	2° ROH One ring	
9 C ₃ H ₆ O	No	No	Yes	No ROH One alkene	

Section 11-5 Conversion of Alcohols to "Tosylates", and their use as Exceptional Leaving Groups in S_N2 , S_N1 , E2, and E1 Reactions

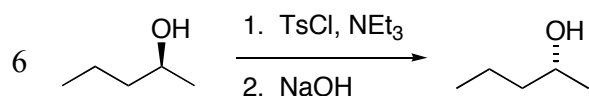
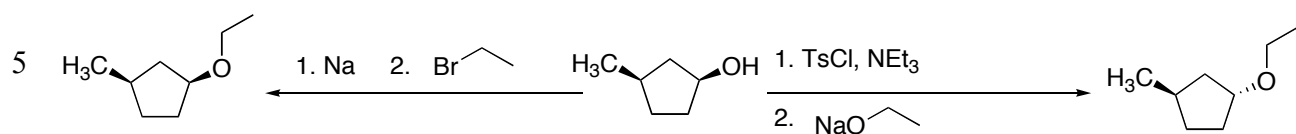
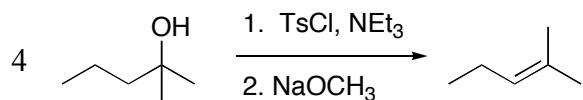
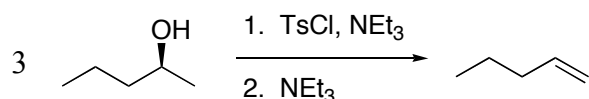
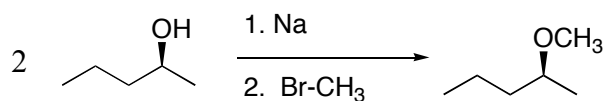
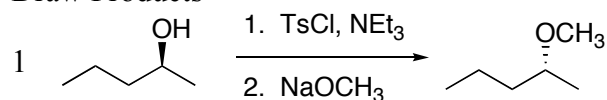
- Tosylates are super leaving groups, better even than iodides.
- Tosylates are well suited to S_N2 and E2 reactions.



Notes:

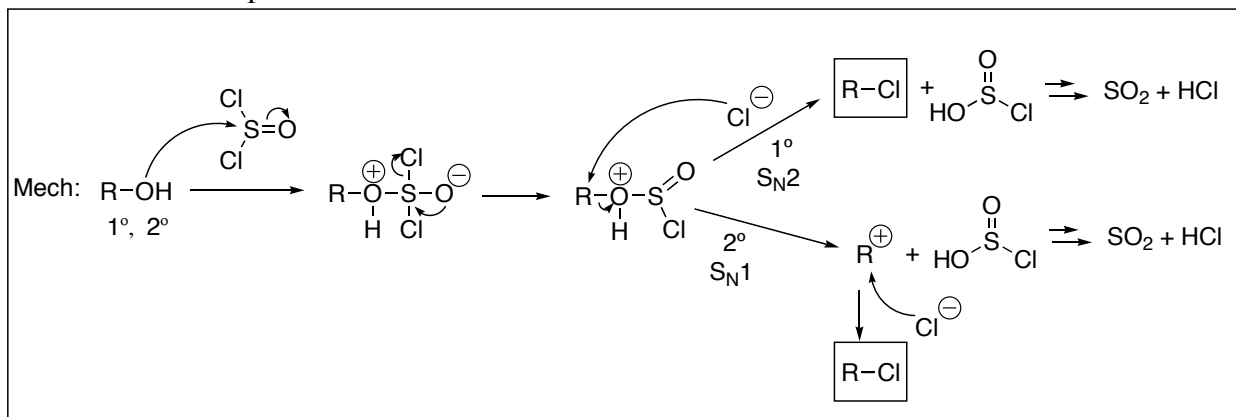
1. Tosylates are easy to form
2. "Toluene sulfonate"
3. Tosylate anion is really stable, comparable to the anion from sulfuric acid
 - Both the electronegative sulfur and the resonance/charge sharing with the other oxygens helps
4. Whereas a normal OH has a poor leaving group (hydroxide anion), conversion to the tosylate provides a super good leaving group.
5. Leaving Group Reactivity: Better than the best of the halides
 - $\text{OTs} \gg \text{I} > \text{Br} > \text{Cl}$
6. Tosylates are highly reactive toward S_N2 , S_N1 , E2, and E1 Reactions
7. Triethylamine is used as an HCl scavenger in the tosylate formation
 - Often a weaker amine base called pyridine is used, to avoid unintentionally providing E2 on the tosylate

Draw Products



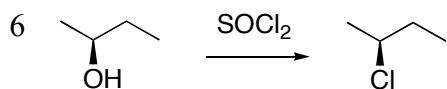
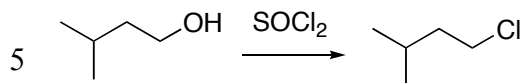
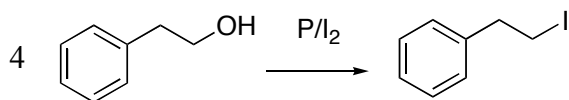
Reaction of 1° and 2° Alcohols with SOCl₂ (Section 11-9)

- Default recipe for 1° and 2° alcohols

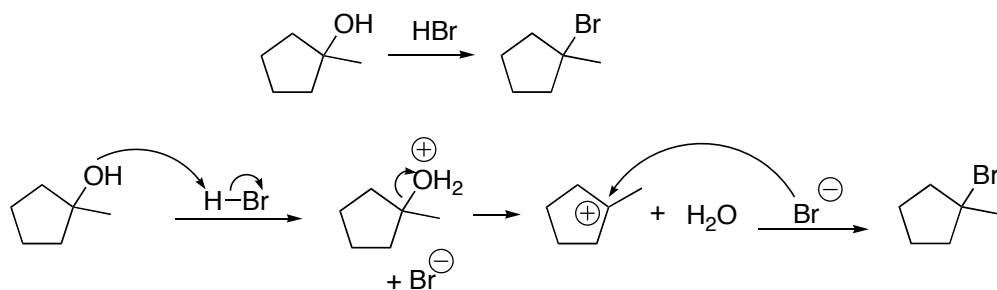


- Mechanism: Not for test responsibility
- Mechanism differs for 1° and 2° alcohols
- 1° involve an S_N2 substitution
- 2° involve an S_N1 type substitution
- The chloride that captures the cation is normally on the same side of the molecule on which the oxygen began, and often captures the cation very rapidly from that same side
- This results in a very unusual **retention of stereochemistry.**
- When they work, these reactions are convenient because the side products, SO₂ and HCl, are both gases. So workup is really easy. Simply rotovap the mixture down, and everything except for product is gone.

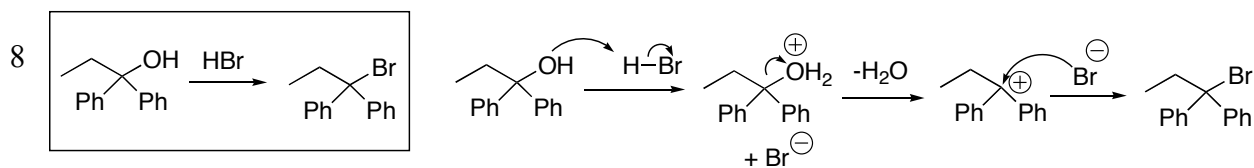
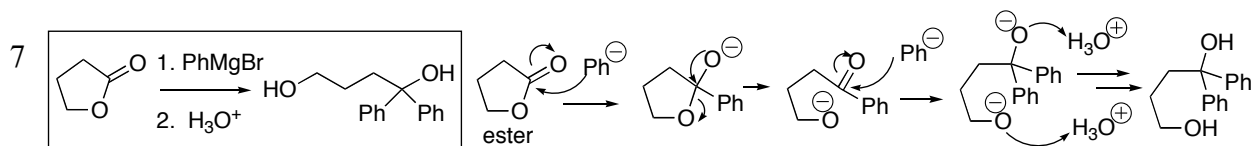
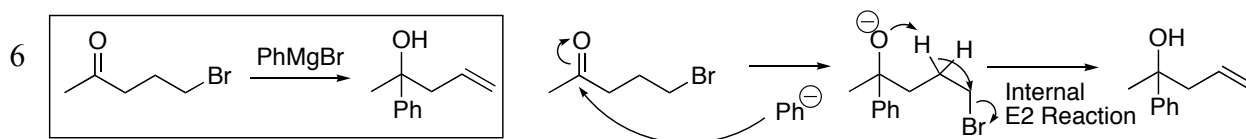
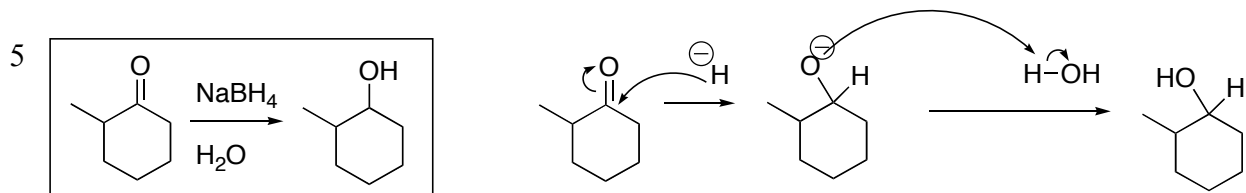
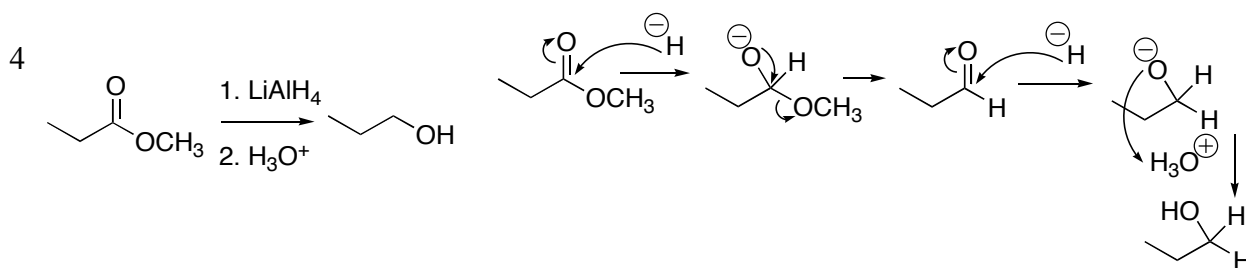
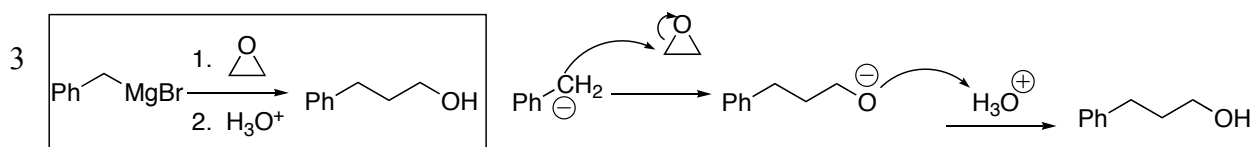
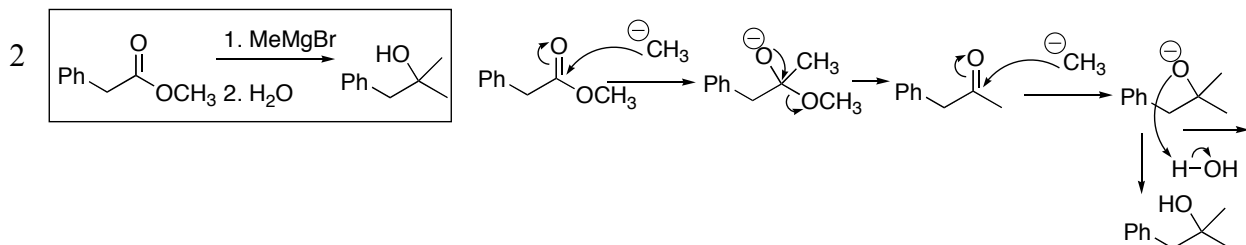
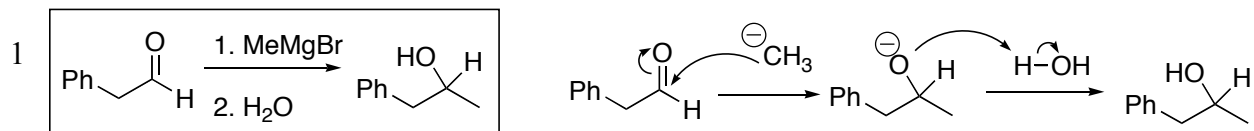
Draw Products or Provide Appropriate Reactants for the following Transformations



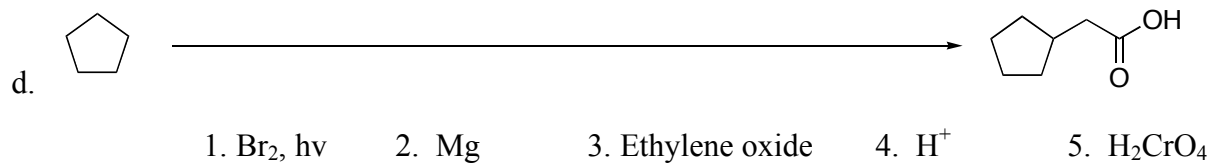
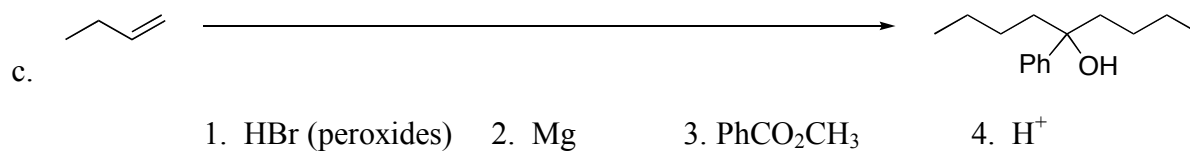
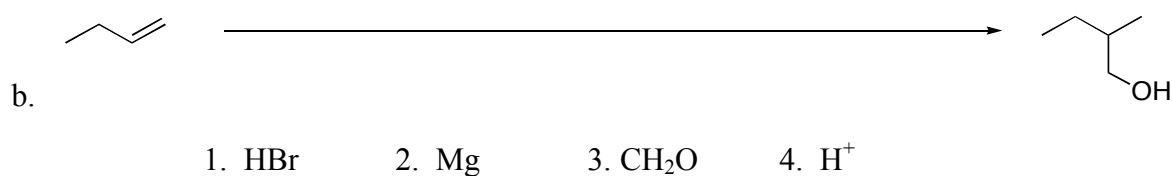
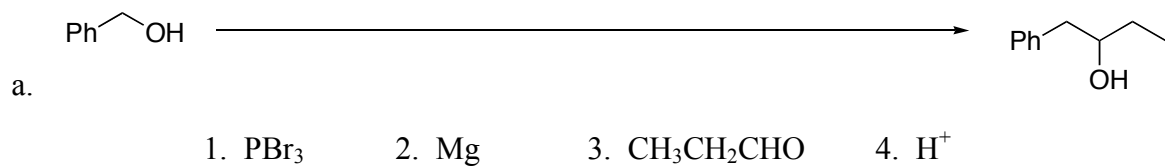
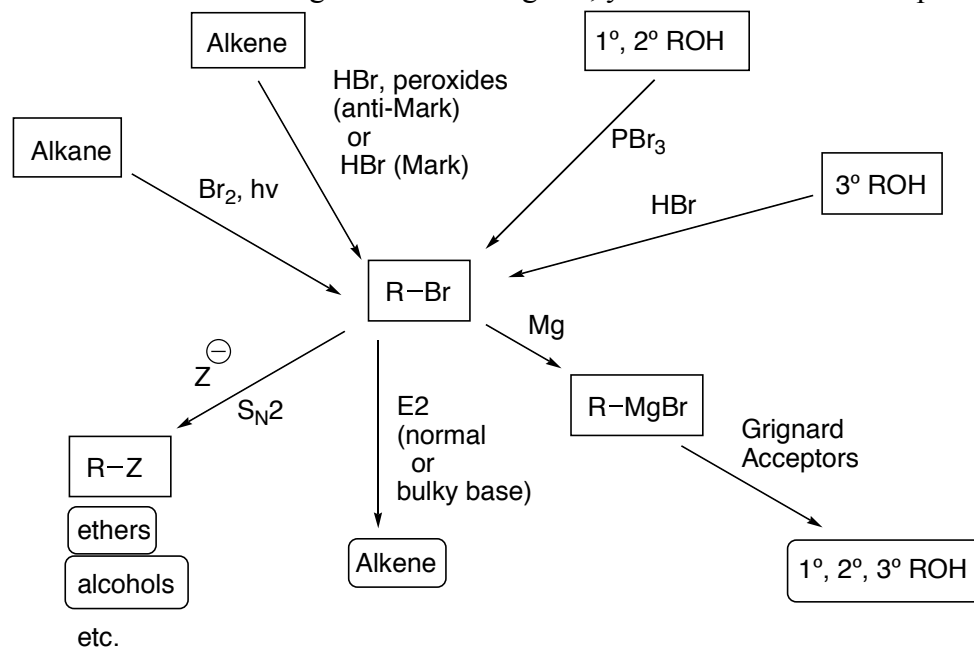
Draw the Mechanism:

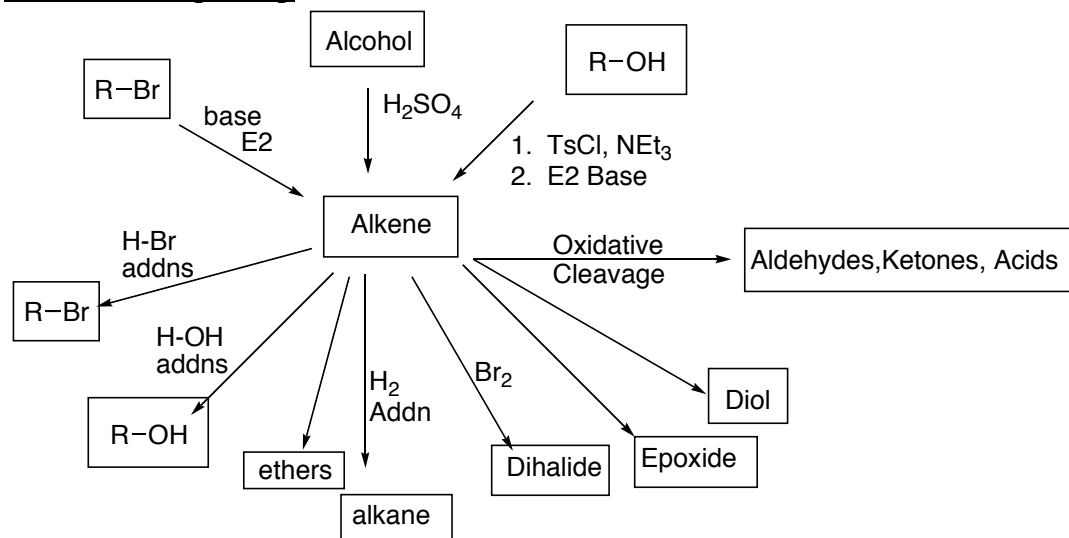


Draw the mechanisms for the following reactions.



REVIEW. To make organometallic reagents, you must have RBr compounds (or RCl or RI).



Alkene Concept Map**Ether Concept Map**