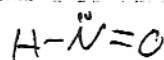
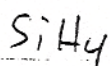
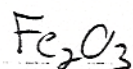
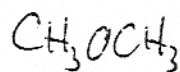
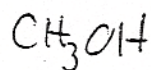
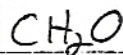


① Identify any Forces (Other than London)
for following:



② Which has London as only binding force?



① which will have highest bp? mp?

SO_2	Cl_2	$\begin{array}{c} C \\ \\ HPOH \end{array}$	$CH_3CH_2CH_2CH_3$
mp = 64	70	65	72

② which has highest mp?
lowest bp?

$NaCl$	$HOCH_2CH_3$	CH_3OCH_3
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Highest bp:

③	OBr_2	OF_2
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④	$CH_3CH_2CH_2NH_2$	$\begin{array}{c} CH_3 \\ \\ CH_3-N-CH_3 \\ \\ CH_3 \end{array}$
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⑤	CH_3OH	CH_3CH_2OH
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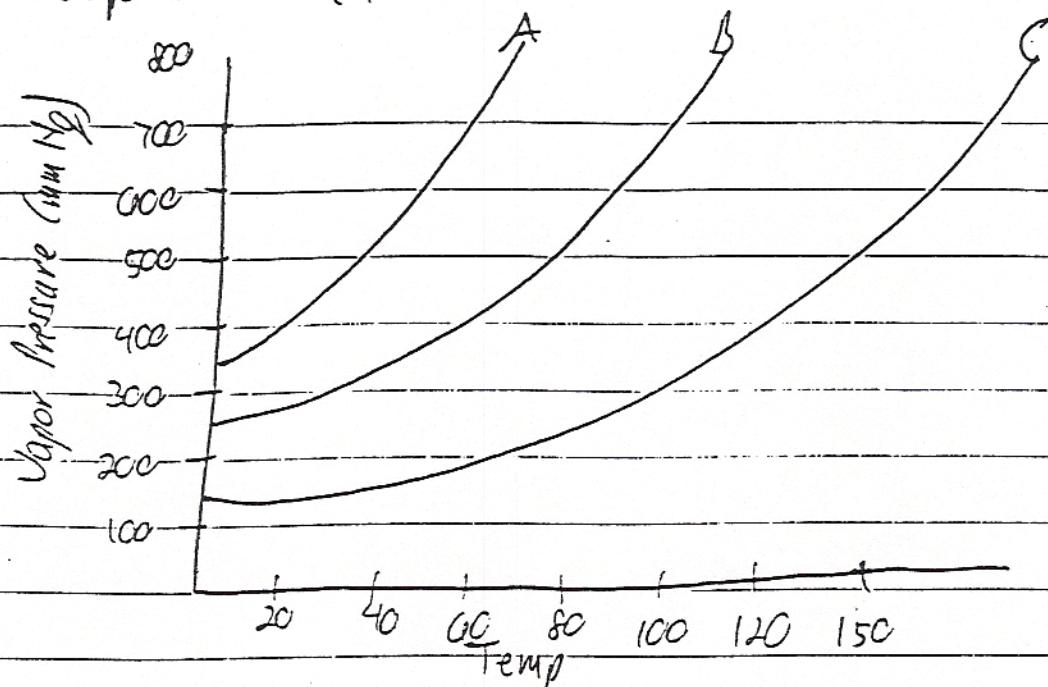
⑥	Br_2	F_2	SiH_4	CO_2
	100	38	32	44

⑦ Rank, from highest to lowest bp?

CH_3NO_2	LiF	Cl_2	CH_3OH
70	76	70	32

Vapor Pressure Curve

(E1-14)



① What's the "normal" bp for B?

② What's the bp for B at 400 mm Hg?

③ What temp needed to get 400 mm Hg for C?

④ Rank the IMF for A, B+C.

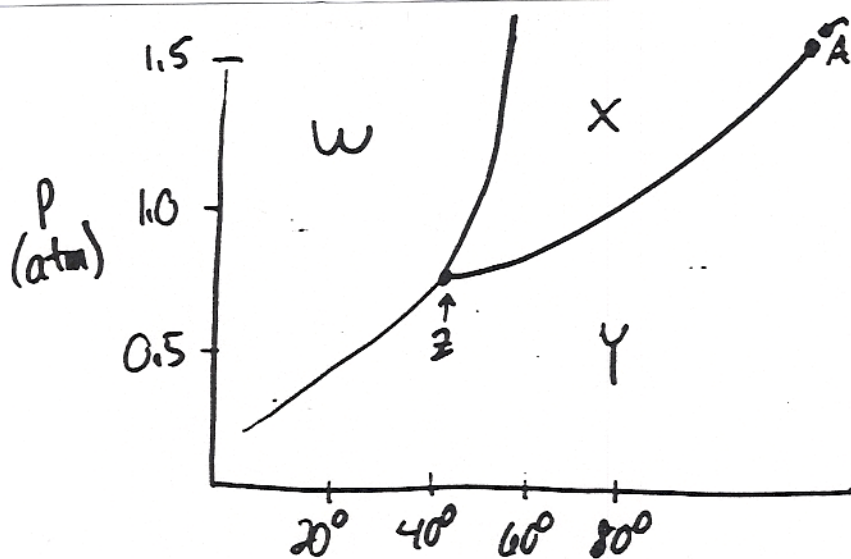
Heating Curve Sample Problem

How much energy (kJ) to heat 36 g ice
(18 g/mol) from $-50^{\circ}\text{C} \rightarrow +50^{\circ}\text{C}$?

Given: $\Delta H_{\text{fusion}} = 6.01 \text{ kJ/mol}$

specific heat (ice) = $2.09 \text{ J/g}\cdot\text{K}$

specific heat (water) = $4.18 \text{ J/g}\cdot\text{K}$



Which letter represents:

- ① Gas Phase
- ② Liquid
- ③ Solid
- ④ Triple Point
- ⑤ Critical Point

⑤ What is the normal bp?

- a) 20° b) 40° c) 60° d) 80°

⑥ What is the normal mp?

- a) 20° b) 50° c) 80°

⑦ When a ~~gas~~ liquid at 60° has pressure adjusted, at what pressure does it vaporize?

⑧ When solid at 0.5 atm is warmed, does it
a) melt or b) sublime